

(19)



Europäisches Patentamt
European Patent Office
Office européen des brevets



(11)

EP 1 193 365 A1

(12)

EUROPEAN PATENT APPLICATION

(43) Date of publication:

03.04.2002 Bulletin 2002/14

(51) Int Cl.7: **E21B 33/138, C09K 7/02**

(21) Application number: 01308209.4

(22) Date of filing: 27.09.2001

(84) Designated Contracting States:

AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU
MC NL PT SE TR

Designated Extension States:

AL LT LV MK RO SI

(30) Priority: 28.09.2000 US 677826

(71) Applicant: Halliburton Energy Services, Inc.
Duncan, Oklahoma 73536 (US)

(72) Inventors:

- Eoff, Larry S.
Duncan, Oklahoma 73533 (US)
- Reddy, Raghava B.
Duncan, Oklahoma 73533 (US)
- Dalrymple, Eldon D.
Duncan, Oklahoma 73533 (US)

(74) Representative: Wain, Christopher Paul et al
A.A. Thornton & Co.
235 High Holborn
London WC1V 7LE (GB)

(54) Reducing subterranean formation water permeability

(57) The water permeability of a subterranean formation is reduced by introducing into the formation a water flow resisting chemical which attaches to adsorption sites on surfaces within the porosity of the formation and reduces the water permeability thereof without substan-

tially reducing the hydrocarbon permeability thereof. The water flow resisting chemical is comprised of a polymer of at least one hydrophilic monomer and at least one hydrophobically modified hydrophilic monomer.

EP 1 193 365 A1

Description

[0001] The present invention relates to a method of treating a water and hydrocarbon producing subterranean formation to reduce the water permeability thereof.

[0002] The production of water with hydrocarbons, i.e. oil and/or gas, from wells constitutes a major problem and expense in the production of the hydrocarbons. While hydrocarbon producing wells are usually completed in hydrocarbon producing formations, when the formations contain layers of water and oil or when there are water producing zones near the hydrocarbon producing formations, the higher mobility of the water often allows it to flow into the wellbores which penetrate the hydrocarbon producing formations by way of natural fractures and/or high permeability streaks. In the production of such wells, the ratios of water to hydrocarbons recovered often become so high that the cost of producing the water, separating it from the hydrocarbons and disposing of it represents a significant economic loss.

[0003] In order to reduce the production of undesired water from hydrocarbon producing formations, aqueous polymer solutions containing cross-linking agents have been utilized heretofore. In the case of naturally fractured formations, such aqueous polymer solutions have been pumped into the hydrocarbon producing formations so that they enter water zones within and adjacent to the formations and cross-link therein. The cross-linking of the polymer solutions causes them to form stiff gels which aid in stopping or reducing the flow of the undesired water. While the use of aqueous polymer solutions for reducing the production of undesired water has achieved varying degrees of success, the full blocking gels produced are not suitable for producing formation treatments unless the polymer solution can be placed solely in the offending water producing zone or zones therein. If a polymer solution is allowed to gel within a hydrocarbon producing zone, the cross-linked polymer gel formed will reduce or stop the flow of hydrocarbons in addition to the flow of water. The selected placement of a polymer solution in a producing formation requires expensive, time-consuming zonal isolation placement technology. In addition, even when a polymer solution is properly placed in a water producing zone, the cross-linked gels formed often do not remain stable in the zone due to thermal degradation and/or differences in the adsorption characteristics of the polymer and associated cross-linker and the like.

[0004] More recently, chemicals referred to as relative permeability modifiers have been utilized to decrease the production of water with hydrocarbons. That is, water permeability modifying chemicals such as polyacrylamide have been introduced into hydrocarbon and water producing formations so that the chemicals attach to adsorption sites on surfaces within the porosity of the formations. The presence of the chemicals in the formations has the effect of reducing the flow of water through the formations while having a minimal effect on the flow of hydrocarbons therethrough. The use of water permeability modifying chemicals in hydrocarbon and water producing formations to decrease the production of water is considerably less expensive than other techniques such as blocking the flow of water with cross-linked polymers, and does not require expensive zonal isolation techniques. However, the use of such hydrophilic water permeability modifying chemicals, e.g. polyacrylamides, have heretofore resulted in only small reductions in water production and/or unacceptable levels of reduction in hydrocarbon production.

[0005] Thus, there are needs for improved methods of treating water and hydrocarbon producing subterranean formations utilizing water permeability modifying chemicals which are more effective in reducing water production with minimal or no reduction in hydrocarbon production.

[0006] In one aspect, the present invention provides a method of treating a water and hydrocarbon producing subterranean formation to reduce the water permeability thereof, which method comprises introducing into said formation a water flow resisting chemical comprised of a polymer of at least one hydrophilic monomer and at least one hydrophobically modified hydrophilic monomer, which polymer attaches to adsorption sites on surfaces within the porosity of said formation to reduce the water permeability thereof without substantially reducing the hydrocarbon permeability.

[0007] Examples of particularly suitable water flow resisting polymers useful in accordance with this invention include, but are not limited to, an acrylamide/octadecyldimethylammoniummethyl methacrylate bromide copolymer, a dimethylaminoethyl methacrylate/vinyl pyrrolidone/hexadecyldimethylammoniummethyl methacrylate bromide terpolymer and an acrylamide/2-acrylamido-2-methyl propane sulfonic acid/2-ethylhexyl methacrylate terpolymer. These and other polymers which can be utilized preferably have a molecular weight in the range of from about 100,000 to about 10,000,000. Further the mole ratio of the hydrophilic monomer(s) to the hydrophobically modified hydrophilic monomer(s) in the polymers is preferably in the range of from about 99.98:0.02 to about 90:10.

[0008] The polymer utilized can be introduced into the formation to be treated in an aqueous carrier liquid solution and overflushed with a water compatible with the formation. After the formation is contacted with the solution and the polymer has attached to surfaces within the porosity of the formation, a secondary overflush of a hydrocarbon such as diesel oil (or a gas in the case of a hydrocarbon gas producing formation) is preferably introduced therein to facilitate the subsequent production of hydrocarbons therethrough.

[0009] The term "water" when used herein in reference to the water produced with hydrocarbons from subterranean formations includes salt water and brines.

[0010] The polymers useful in accordance with this invention can be prepared from a variety of hydrophilic monomers

and hydrophobically modified hydrophilic monomers.

[0011] Examples of particularly suitable hydrophilic monomers which can be utilized include, but are not limited to, acrylamide, 2-acrylamido-2-methyl propane sulfonic acid, N,N-dimethylacrylamide, vinyl pyrrolidone, dimethylaminoethyl methacrylate, acrylic acid, dimethylaminopropylmethacrylamide, trimethylammoniummethyl methacrylate chloride, methacrylamide and hydroxyethyl acrylate. Of these, acrylamide, 2-acrylamido-2-methyl propane sulfonic acid, acrylic acid, dimethylaminoethyl methacrylate and vinyl pyrrolidone are preferred.

[0012] A variety of hydrophobically modified hydrophilic monomers can also be utilized to form the polymers useful in accordance with this invention. Particularly suitable hydrophobically modified hydrophilic monomers include, but are not limited to, alkyl acrylates, alkyl methacrylates, alkyl acrylamides and alkyl methacrylamides wherein the alkyl radicals have from about 4 to about 22 carbon atoms, alkyl dimethylammoniummethyl methacrylate bromide, alkyl dimethylammoniummethyl methacrylate chloride and alkyl dimethylammoniummethyl methacrylate iodide wherein the alkyl radicals have from about 4 to about 22 carbon atoms and alkyl dimethylammonium-propylmethacrylamide iodide wherein the alkyl groups have from about 4 to about 22 carbon atoms. Of these, octadecyldimethylammoniummethyl methacrylate bromide, hexadecyldimethylammoniummethyl methacrylate bromide, hexadecyldimethylammoniumpropyl methacrylamide bromide, 2-ethylhexyl methacrylate and hexadecyl methacrylamide are preferred.

[0013] Polymers which are useful in accordance with the present invention can be prepared by polymerizing any one or more of the hydrophilic monomers with any one or more of the hydrophobically modified hydrophilic monomers. While the polymerization reaction can be performed in various ways, an example of a particularly suitable procedure for polymerizing water soluble monomers is as follows. Into a 250 mL-3 neck round bottom flask, charge the following: 47.7 g DI water, 1.1 g acrylamide and 0.38 g alkyl dimethylammoniummethyl methacrylate bromide. The solution formed is sparged with nitrogen for approximately 30 minutes, followed by the addition of 0.0127 g of 2,2'-azobis (2-amidinopropane) dihydrochloride. The resulting solution is then heated, with stirring, to 110°F and held for 18 hours to produce a highly viscous polymer solution.

[0014] When the hydrophobically modified hydrophilic monomer is not water soluble, e.g., octadecylmethacrylate, the following procedure can be utilized. Into a 250 mL-3 neck round bottom flask, charge the following: 41.2 g DI water and 1.26 g acrylamide. The solution formed is sparged with nitrogen for approximately 30 minutes, followed by the addition of 0.06 g of octadecyl methacrylate and 0.45 g of a cocoamidopropyl betaine surfactant. The mixture is stirred until a homogeneous, clear solution is obtained followed by the addition of 0.0055 g of 2,2'-azobis (2-amidinopropane) dihydrochloride. The resulting solution is then heated, with stirring, to 110°F and held for 18 hours to produce a highly viscous polymer solution.

[0015] In addition, the polymerization procedure may employ a hydrocarbon reaction medium instead of water. In this case, appropriate surfactants are used to emulsify the hydrophilic/hydrophobic monomers, and the product is obtained as an oil external/water internal emulsion.

[0016] Suitable polymers prepared as described above have estimated molecular weights in the range of from about 250,000 to about 3,000,000 and have mole ratios of the hydrophilic monomer(s) to the hydrophobically modified hydrophilic monomer(s) in the range of from about 99.98:0.02 to about 90:10. Particularly suitable polymers having molecular weights and mole ratios in the ranges set forth above include, but are not limited to, an acrylamide/octadecyldimethylammoniummethyl methacrylate bromide copolymer, a dimethylaminoethyl methacrylate/hexadecyldimethylammoniummethyl methacrylate bromide copolymer, a dimethylaminoethyl methacrylate/vinyl pyrrolidone/hexadecyldimethylammoniummethyl methacrylate bromide terpolymer and an acrylamide/2-acrylamido-2-methyl propane sulfonic acid/2-ethylhexyl methacrylate terpolymer. Of these, an acrylamide/octadecyl dimethylammoniummethyl methacrylate bromide copolymer having a mole ratio of hydrophilic monomer to hydrophobically modified hydrophilic monomer of 96:4 is presently preferred.

[0017] An improved method of this invention for treating a water and hydrocarbon producing subterranean formation to reduce the water permeability thereof without substantially reducing the hydrocarbon permeability thereof is comprised of the following steps. A polymer of at least one hydrophilic monomer and at least one hydrophobically modified hydrophilic monomer is prepared having a molecular weight in the range of from about 100,000 to about 10,000,000 and having a mole ratio of hydrophilic monomer(s) to hydrophobically modified hydrophilic monomer(s) in the range of from about 99.98:0.02 to about 90:10. Thereafter, the polymer is introduced into the subterranean formation to be treated so that the polymer attaches to adsorption sites on surfaces within the porosity of the formation.

[0018] In carrying out the above described method, the polymer is preferably dissolved in an aqueous carrier liquid and the resulting solution containing the polymer is introduced into the formation. The aqueous carrier liquid can be fresh water, seawater, or an aqueous salt solution. The aqueous carrier liquid is preferably an aqueous salt solution containing one or more salts in an amount in the range of from about 2% to about 10% by weight of the solution. A variety of salts can be utilized in the aqueous solution such as potassium chloride, sodium chloride, ammonium chloride and calcium chloride.

[0019] In order to facilitate the flow of the carrier liquid solution containing the polymer into the formation being

treated, a surfactant can be included in the carrier fluid solution. While a variety of surfactants can be used, a presently preferred surfactant is cocoylamidopropylbetaine. When used, the surfactant is included in the carrier liquid solution in an amount in the range of from about 0.10% to about 2.0% by weight of the solution, more preferably in an amount in the range of from about 0.5% to about 1%.

[0020] After the treatment described above whereby the polymer utilized is adsorbed onto the formation surfaces, an after-flush of a hydrocarbon liquid such as kerosene, diesel oil or crude oil or a hydrocarbon or inert gas such as methane and natural gas or nitrogen (when the formation produces gas) can be introduced into the formation. Although it is not required in order for the polymer to be effective, the hydrocarbon liquid or gas after-flush in the formation facilitates the subsequent flow of hydrocarbons through the formation.

[0021] In order to further illustrate the methods of the present invention, the following examples are given.

Example 1

[0022] Various cationic hydrophobically modified hydrophilic monomers were prepared in accordance with the procedures described above by reacting dimethylaminoethyl methacrylate with hexadecyl bromide, octadecyl bromide and 1-bromodocosane (behenyl bromide) to produce hexadecyl, octadecyl and behenyl dimethylammonium-ethyl methacrylate bromide. Various mole percentages of the hydrophobically modified hydrophilic monomers and acrylamide were then reacted to produce copolymers having estimated molecular weights in the range of from about 250,000 to about 3,000,000. In addition, a non-ionic hydrophobic monomer, i.e., octadecylmethacrylate was dissolved in an aqueous solution containing a suitable surfactant and copolymerized with acrylamide as described above. The various copolymers produced are described in TABLE I below.

[0023] For comparison purposes, polyacrylamide and two commercially available polyacrylamide polymers of known molecular weight, i.e., "ALCOFLOOD 245S™" and "ALCOFLOOD 935™" are included in TABLE I.

TABLE I
Copolymers Tested

Polymer or Copolymer Number	Copolymer Description		Hydrophilic Monomer, %	Mole % Ratio Hydrophobically Modified Hydrophilic Monomer, %	Viscosity ¹ , cp
	Hydrophilic Monomer	Hydrophobically Modified Hydrophilic Monomer			
1	(Control - polyacrylamide)		-	-	5.7
1(a)	"ALCOFLOOD 245ST" ¹		-	-	1.8
1(b)	"ALCOFLOOD 935T" ¹		-	-	29
2	acrylamide / hexadecyl dimethylammonium ethylmethacrylate bromide		98	2	-
3	acrylamide / hexadecyl dimethylammonium ethylmethacrylate bromide		99	1	-
4	acrylamide / octadecyl dimethylammonium ethylmethacrylate bromide		97	3	-
5	acrylamide / octadecyl dimethylammonium ethylmethacrylate bromide		98	2	-
6	acrylamide / octadecyl dimethylammonium ethylmethacrylate bromide		99	1	-
7	acrylamide / octadecyl dimethylammonium ethylmethacrylate bromide		99.5	0.5	-
8	acrylamide / octadecyl dimethylammonium ethylmethacrylate bromide		99.98	0.02	-
9	acrylamide / octadecyl methacrylate		99	1	-
10	acrylamide / octadecyl dimethylammonium ethylmethacrylate bromide		98	2	-
11	acrylamide / octadecyl dimethylammonium ethylmethacrylate bromide		99	1	-
12	acrylamide / hexadecyl dimethylammonium ethylmethacrylate bromide		96	4	-
13	acrylamide / hexadecyl dimethylammonium ethylmethacrylate bromide		95	5	-
14	acrylamide / octadecyl dimethylammonium ethylmethacrylate bromide		97	3	-
15	acrylamide / octadecyl dimethylammonium ethylmethacrylate bromide		96	4	-

¹ "ALCOFLOOD 245ST"¹ is commercially available from Ciba Specialty Chemicals of Suffolk, Virginia, and has a molecular weight of 500,000.

1 "ALCOFLOOD 935™" is commercially available from Ciba Specialty Chemicals of Suffolk, Virginia, and has a molecular weight of 5 to 6 million.
 2 The viscosities were run at 0.4% active polymer by weight in a 6% by weight aqueous NaCl solution using a Brookfield LVT viscometer with a UL
 3 adaptor at 12 rpm for "ALCOFLOOD 935™" and at 60 rpm for Polymer No. 1 and "ALCOFLOOD 245™".

[0024] A multi-pressure tap Hassler sleeve containing a Berea sandstone core was utilized to determine the water permeability reduction produced by the various copolymers described in Table I. The permeability reduction tests were run at a temperature of 175°F utilizing a brine containing 9% by weight sodium chloride and 1% by weight calcium

chloride.

[0025] The following procedure was utilized for a first series of tests, the results of which are shown in Table II below. The above described brine was flowed through the Berea core until the pressure stabilized, yielding an initial permeability. The polymer or copolymer tested was dissolved in the above described brine in an amount of 2,000 parts per million, and 100 milliliters of the treatment solution containing the polymer or copolymer was flowed into the core and overflushed through the core with 100 milliliters of the above described brine, both of which were flowed in the opposite direction to the initial brine flow. The brine flow was then resumed in the original direction until the pressure again stabilized and final permeability was calculated. The percent brine permeability reduction was calculated using the formula $[1 - (\text{final permeability} / \text{initial permeability})] \times 100$. The multi-tap Hassler sleeve allowed the core permeability to be divided into four segments. In the tests, the initial brine flow was from segment 1 through segment 4. The treatment solution and displacement brine flow were from segment 4 through segment 1, and the final brine flow was from segment 1 through segment 4. The initial and final permeabilities were calculated for the middle two segments, i.e., segments 2 and 3. The results of the tests are set forth in Table II below.

TABLE II

Permeability Reduction in Berea Sandstone At 175°		
Polymer or Copolymer No.	% Brine Permeability Reduction - Segment 2	% Brine Permeability Reduction - Segment 3
1 (Control)	38	20
2	33	53
3	35	35
4	30	43
5	43	52
6	48	65
7	10	42
8	60	60
9	45	35

[0026] From Table II, it can be seen that the copolymers of this invention produced greater brine reduction than polymer No. 1 (unmodified polyacrylamide).

Example 2

[0027] The tests described in Example 1 were repeated except that a different flow sequence was utilized to simulate a subterranean zone which initially produced oil but later "watered out," i.e., produced only brine but still had residual oil trapped therein. This scenario was simulated by flowing a sequence of brine and oil (kerosene) through the core. In a first set of tests, the sequence was first brine, then oil, then brine, then oil, then brine, then the treatment solution and then brine. The third brine flow was used to calculate the initial permeability to brine at residual oil conditions. This permeability was used along with the permeability calculated after the final brine flow to determine the percent brine permeability reduction. To determine the impact of thermal degradation, following the final brine flow, some testing was performed where the core was shut-in at temperature over night and flow was resumed the next day. The brine permeability was determined after the shut-in.

[0028] Another series of tests were run to simulate what happens if the treatment solution enters an oil-producing zone containing residual water. These tests were conducted with a flow sequence of a first brine flow, then an oil flow, then a second brine flow, then a second oil flow, then the treatment solution flow followed by a third oil flow. The second oil flow was used to calculate the initial oil permeability and the last oil flow was used to determine the permeability to oil after treatment. The percent oil permeability reduction was determined in the same manner as described above for determining the percent brine permeability reduction. The results of these tests are given in Table III below.

TABLE III

Permeability Reduction In Berea Sandstone At 175°F						
Polymer or Copolymer No.	% Brine Permeability Reduction -Segment 2	% Brine Permeability Reduction -Segment 3	% Brine Permeability Reduction -Segment 2 24 Hour Shut-in	% Brine Permeability Reduction -Segment 3 24 Hour Shut-in	% Oil Permeability Reduction -Segment 2	% Oil Permeability Reduction -Segment 3
1 (Control)	25	45	-65	-30	-	-
3	45	55	-	-	-30	0
4	55	72	-	-	20	-10
5	45	70	-	-	-50	-40
10	35	50	35	50	-	-
11	55	75	-	-	-	-

[0029] From Table III, it can be seen that the copolymers of this invention functioned to increase the brine permeability reduction as compared to polymer No. 1 (polyacrylamide).

[0030] It can also be seen that polymer No. 1 (polyacrylamide) lost effectiveness after 24 hours. In comparison, copolymer No. 10 retained its effectiveness after the 24 hour shut-in. The percent oil permeability reduction numbers were negative indicating increased oil permeability after the treatment.

Example 3

[0031] The tests described in Examples 1 and 2 (with the exception of the oil permeability reduction tests of Example 2) were repeated except that a different flow sequence was utilized. That is, the flow sequence was first brine, then oil, then brine. The second brine flow was used to calculate the initial permeability to brine under residual oil conditions. The treatment solution was then flowed through the core and overflushed through the core with 100 milliliters of brine which was followed by 25 milliliters of oil (kerosene). Following the treatment and overflush stages, brine flow was again stabilized through the core in the original direction. The percent brine permeability reduction in segments 2 and 3 before and after a 24 hour shut-in were calculated. The results of these tests are given in Table IV below.

TABLE IV

Permeability Reduction In Berea Sandstone At 175°F				
Polymer or Copolymer No.	% Brine Permeability Reduction - Segment 2	% Brine Permeability Reduction - Segment 3	% Brine Permeability Reduction - Segment 2 24 Hour Shut-in	% Brine Permeability Reduction - Segment 3 24 Hour Shut-in
1	40	63	0	17
12	42	68	32	60
13	65	85	55	80
10	72	78	58	68
14	65	82	48	74
15	78	83	78	83

[0032] From Table IV, it can be seen that the control polyacrylamide lost its effectiveness after the 24 hour shut-in. The copolymers of the present invention produced higher levels of brine permeability reduction and maintained their effectiveness after the 24 hour shut-in.

Claims

1. A method of treating a water and hydrocarbon producing subterranean formation to reduce the water permeability thereof, which method comprises introducing into said formation a water flow resisting chemical comprised of a polymer of at least one hydrophilic monomer and at least one hydrophobically modified hydrophilic monomer, which polymer attaches to adsorption sites on surfaces within the porosity of said formation to reduce the water permeability thereof without substantially reducing the hydrocarbon permeability.
2. A method according to claim 1, wherein said hydrophilic monomer is selected from acrylamide, 2-acrylamido-2-methyl propane sulfonic acid, N,N-dimethylacrylamide, vinyl pyrrolidone, dimethylaminoethyl methacrylate, acrylic acid, dimethylaminopropylmethacrylamide, trimethylammoniummethyl methacrylate chloride, methacrylamide and hydroxyethyl acrylate.
3. A method according to claim 1 or 2, wherein said hydrophobically modified hydrophilic monomer is selected from alkyl acrylates, alkyl methacrylates, alkyl acrylamides and alkyl methacrylamides wherein the alkyl radicals have from 4 to 22 carbon atoms; alkyl dimethylammoniummethyl methacrylate bromide, alkyl dimethylammoniummethyl methacrylate chloride and alkyl dimethylammoniummethyl methacrylate iodide wherein the alkyl radicals have from 4 to 22 carbon atoms; and alkyl dimethylammonium propylmethacrylamide bromide, alkyl dimethylammonium propylmethacrylamide chloride and alkyl dimethylammonium propylmethacrylamide iodide wherein the alkyl groups have from 4 to 22 carbon atoms.
4. A method according to claim 1, 2 or 3, wherein the mole ratio of said hydrophilic monomer to said hydrophobically modified hydrophilic monomer in said polymer is from 99.98:0.02 to 90:10.
5. A method according to claim 1, 2, 3 or 4, wherein said polymer is selected from an acrylamide/octadecyldimethylammoniummethyl methacrylate bromide copolymer, a dimethylaminoethyl methacrylate/hexadecyldimethylammoniummethyl methacrylate bromide copolymer, a dimethylaminoethyl methacrylate/vinyl pyrrolidone/ hexadecyldimethylammoniummethyl methacrylate bromide terpolymer and an acrylamide/2-acrylamido-2-methyl propane sulfonic acid/2-ethylhexyl methacrylate terpolymer.
6. A method according to any of claims 1 to 5, wherein said polymer has a molecular weight in the range of 100,000 to 10,000,000.
7. A method according to any of claims 1 to 6, wherein said polymer is dissolved in an aqueous carrier liquid, preferably an aqueous salt solution.
8. A method according to claim 7, wherein the carrier liquid is an aqueous salt solution selected from potassium chloride, sodium chloride, ammonium chloride and calcium chloride, the salt being present in said aqueous salt solution in an amount of from 2% to 10% by weight of said solution.
9. A method according to claim 7 or 8, wherein said aqueous carrier liquid solution further comprises a surfactant dissolved therein, preferably cocoylamidopropylbetaine in an amount of from 0.1% to 2% by weight of said carrier liquid solution.
10. A method according to any of claims 1 to 9, which further comprises the step of introducing a hydrocarbon liquid or a gas into said formation after the introduction of said water flow resisting chemical therein, said hydrocarbon liquid preferably being kerosene, diesel oil or crude oil, and said gas preferably being methane, natural gas or nitrogen.

European Patent
Office

EUROPEAN SEARCH REPORT

Application Number
EP 01 30 8209

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (MCL7)
X	WO 93 15164 A (SCHLUMBERGER TECHNOLOGY CORP ; SCHLUMBERGER SERVICES PETROL (FR); I) 5 August 1993 (1993-08-05) * page 2, line 16 - page 3, line 34 * * page 4, line 26 - page 5, line 32 *	1-3	E21B33/138 C09K7/02
Y	US 4 532 052 A (WEAVER JIMMIE D ET AL) 30 July 1985 (1985-07-30) * column 5, line 53 - column 9, line 63 * * column 17, line 4 - column 18, line 4 * * column 19, line 7 - column 20, line 29 * * column 22, line 37 - line 46 *	1-3,7	
Y	US 4 604 216 A (IRVIN HOWARD B ET AL) 5 August 1986 (1986-08-05) * column 1, line 61 - column 2, line 30 * * column 3, line 57 - column 4, line 2 * * column 4, line 60 - line 68; claims 1-5,9,13 *	1-3,6-8	
A		4	
Y	GB 2 128 659 A (PHILLIPS PETROLEUM CO) 2 May 1984 (1984-05-02) * page 1, line 30 - line 48 * * page 2, line 5 - line 56; claims 1-5 *	1-3,6-8	TECHNICAL FIELDS SEARCHED (MCL7)
A		4	E21B C09K
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 15 January 2002	Examiner Boulon, A
CATEGORY OF CITED DOCUMENTS X: particularly relevant if taken alone Y: particularly relevant if combined with another document of the same category A: technological background O: non-written disclosure P: intermediate document T: theory or principle underlying the invention E: earlier patent document, but published on, or after the filing date D: document cited in the application L: document cited for other reasons &: member of the same patent family, corresponding document			

EPO FORM 1503 03 02 (P0401)

**ANNEX TO THE EUROPEAN SEARCH REPORT
ON EUROPEAN PATENT APPLICATION NO.**

EP 01 30 8209

This annex lists the patent family members relating to the patent documents cited in the above-mentioned European search report. The members are as contained in the European Patent Office EDP file on
The European Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

15-01-2002

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
WO 9315164	A	05-08-1993	FR 2686892 A1	06-08-1993
			DE 69306674 D1	30-01-1997
			DE 69306674 T2	30-04-1997
			DK 578806 T3	16-06-1997
			EP 0578806 A1	19-01-1994
			WO 9315164 A1	05-08-1993
			US 5677266 A	14-10-1997
			US 5759962 A	02-06-1998
			US 5728653 A	17-03-1998
			US 5972848 A	26-10-1999
US 4532052	A	30-07-1985	AU 5116979 A	03-04-1980
			BR 7906202 A	04-11-1980
			CA 1155988 A1	25-10-1983
			DE 2939114 A1	10-04-1980
			GB 2037845 A ,B	16-07-1980
			IT 1124566 B	07-05-1986
			NL 7904632 A	01-04-1980
			NO 793108 A ,B,	31-03-1980
			US 4460627 A	17-07-1984
US 4604216	A	05-08-1986	NONE	
GB 2128659	A	02-05-1984	AU 2010183 A	03-05-1984
			CA 1259182 A1	12-09-1989
			GB 2173507 A ,B	15-10-1986
			NL 8303610 A	16-05-1984
			NO 833775 A ,B,	24-04-1984

EPO FORM P4/89

For more details about this annex : see Official Journal of the European Patent Office, No. 12/82

1000 foreign Opn. : - 6

THIS PAGE BLANK (USPTO)